Temperature: "energy" as a reactant or a product.

$A + B + \text{energy} \leftrightarrow C$

[A] ↑ eq. shifts to right

T° ↑ means [energy] ↑ eq shifts to right

ex: $X + Y \leftrightarrow Z + W \quad \Delta H = -50\,kJ$ (↑ e+o)

stress: T° decreases

$X + Y \rightarrow Z + W + \text{energy}$

T° ↓ means [energy] ↓

[Product] ↓

according to Le Chatelier, system will oppose the change by making more products ∴ shifts to the right.

(mol/L)	A	+ B	⇀ C
I	✓	✓	0
C			
E	?		

$K_{eq} = ✓$ or $[\,]_{eq} = ✓$

$2CO_2 \rightarrow 2CO + O_2$ (g) (g) (g)

$K_{eq} = \dfrac{[CO]^2[O_2]^1}{[CO_2]^2}$

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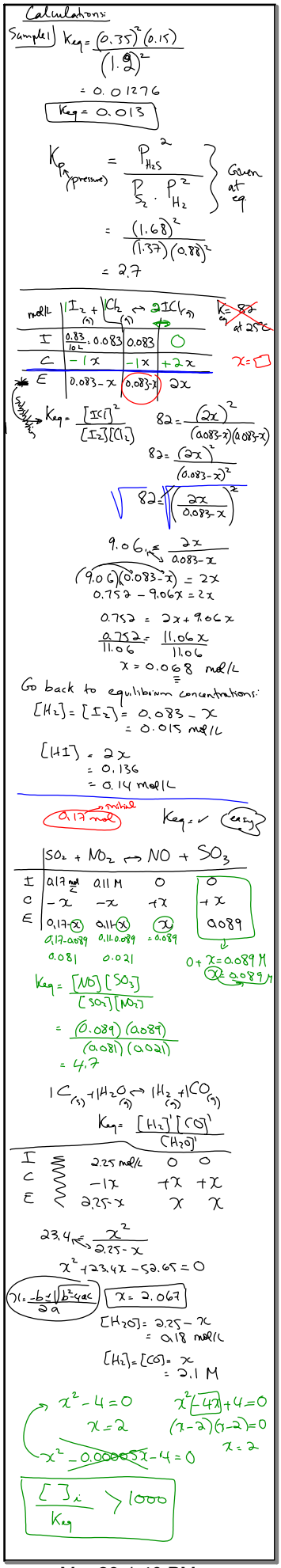
Calculations:
Sample 1
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